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CHAPTER 18: ELECTROCHEMISTRY 677
18.113 Viewed externally, the anode

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looks negative because of the flow of the electrons (from $\text{Zn} \rightarrow \text{Zn}^{2+} + 2\text{e}^{-}$) toward the cathode. In solution, anions move toward the anode because they are attracted by the Zn^{2+} ions surrounding the anode. 18.114 From Table 18.1 of the text.

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If a solution containing A^{+} ions and a solution containing B^{-} ions has just been mixed and $[\text{A}^{+}][\text{B}^{-}]$ is greater than K_{sp} of AB , precipitation occurs. AgI has a K_{sp} of 8.3×10^{-17} , ZnS has a K_{sp} of 1.6×10^{-24} , PbS has a K_{sp} of 8.0×10^{-28} , and CaSO_4 has a K_{sp} of 9.1×10^{-6} .

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Chapter 18 - Reaction Rates &
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idea of reversible reactions and their
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emphasis is placed on how equilibrium
and the rate of a reaction can be
affected by altering temperature,
pressure and concentrations.

Chapter 18 - Reaction Rates & Equilibrium - Mrs. Gingras ...

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Explain why each of the following structures (as shown) is ...

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Unit 1 - Chapters 1 & 2 - Mrs. Gingras' Chemistry Page

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16, Solutions (continued) 7 Look at Figure
162 on page 472 Circle the letter of each
sentence that is true about a saturated
solution a The total amount of dissolved
solute remains constant b The total
mass of [DOC] Chapter 18 Solutions
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Employing Hückel's Rule, label all aromatic species in ...

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